Substance	Ag_2MoO_4(cr)
Third Law Entropy of Silver Molybdate, M.Morishita, H. Houshiyama, Y. Kinoshita, Ai. Nozaki and H. Yamamoto, Mater. Trans., **58**, (2017), pp. 868-872.	
S°_m(Ag_2MoO_4(cr), 298.15 K) / J K^{-1} mol^{-1} = 219.87 ± 2.20 $\Delta_f G^\circ_m$(Ag_2MoO_4(cr), 298.15 K) / kJ mol^{-1} = − 747.49 ± 2.11	

Substance	$BaMoO_4$(cr)
Thermodynamic Properties of Molybdate ion: Reaction Cycles and Experiments, H. Gamjäger and M.Morishita, Pure and Applied Chemistry, **87**, (2015), pp. 461-476.	
$\Delta_{sln} G^\circ_m$($BaMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 48.51 ± 0.22	

Substance	$BaMoO_4$(cr)
Third Law Entropy of Barium Molybdate, M.Morishita, M. Fukushima and H. Houshiyama, Mater. Trans., **57** (2016), pp.46-51.	
S°_m($BaMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 152.69 ± 1.53 Θ_D($BaMoO_4$(cr)) / K = 295 ± 3 $\Delta_f G^\circ_m$($BaMoO_4$, 298.15 K) / kJ mol^{-1} = −1443.29 ± 1.33	

Substance	$BaMoO_4$(cr)
Thermodynamic Properties for $MMoO_4$ (M = Mg, Sr and Ba) as the End-members of the Yellow Phases Formed in the Nuclear Fuel Waste Glasses, M. Morishita, Y. Kinoshita, A. Nozaki and H. Yamamoto Appl. Geo-chem., **98**, (2018),pp.310-320.	
S°_m($BaMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 152.61 ± 1.53 $\Delta_f G^\circ_m$($BaMoO_4$, 298.15 K) / kJ mol^{-1} = −1443.27 ± 1.33	

Substance	$CaMoO_4$(cr)
Thermodynamic Properties of Molybdate ion: Reaction Cycles and Experiments, H. Gamjäger and M.Morishita, Pure and Applied Chemistry, **87**, (2015), pp. 461-476.	
$\Delta_{sln} G^\circ_m$($CaMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 45.69 ± 0.32	

Substance $CaMoO_4$(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.
S°_m($CaMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 122.23 ± 1.22 $\Delta_f G^\circ_m$($CaMoO_4$(cr), 298.15 K) / kJ mol^{-1} = −1437.78 ± 1.11

Substance $CaMoO_4$(cr)
Thermodynamic Properties of Molybdate ion: Reaction Cycles and Experiments, H. Gamjäger and M.Morishita, Pure and Applied Chemistry, **87**, (2015), pp. 461-476.
$\Delta_{sln} G^\circ_m$($CaMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 45.69 ± 0.32

Substance $Ce_2(MoO_4)_3$(cr)
Thermodynamic Properties of Cerium Molybdate, A. Nozaki, M. Morishita, Y. Kinoshita and H. Yamamoto, Z. Metallkd. /Mater. Res. Adv. Tech., **110**, (2019), pp. 715-725.
S°_m($Ce_2(MoO_4)_3$(cr), 298.15 K) / J K^{-1} mol^{-1} = 408.89 ± 4.09 $\Delta_f G^\circ_m$($Ce_2(MoO_4)_3$(cr), 298.15 K) / kJ mol^{-1} = − 4076.64 ± 4.00 T_N ($Ce_2(MoO_4)_3$(cr)) / K = 1.15 ± 0.30 $\Delta_{sln} G^\circ_m$($Sm_2(MoO_4)_3$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 74.25 ± 1.65

Substance $CoMoO_4$(cr)
Redetermination of Standard Gibbs Energiers of Formation of $CoMoO_4$ and $Co_2Mo_3O_8$ by Electromotive Force Measurement, K.Koyama and N.Maekawa, J. Japan Inst. Met., Vol.61 (1997), pp.135-139.
$\Delta_f G^\circ_m$($CoMoO_4$(cr), 1119 - 1273 K) / kJ mol^{-1} = − 1011 + 0.3044 T ±2

Substance $Co_2Mo_3O_8$(cr)
Redetermination of Standard Gibbs Energiers of Formation of $CoMoO_4$ and $Co_2Mo_3O_8$ by Electromotive Force Measurement, K.Koyama and N.Maekawa, J. Japan Inst. Met., Vol.61 (1997), pp.135-139.
$\Delta_f G^\circ_m$($Co_2Mo_3O_8$(cr), 1119 - 1273 K) / kJ mol^{-1} = − 2285 + 0.7000 T ±5

Substance	$FeMoO_4$(cr)
Determination of Standard Gibbs Energies of Formation of $Fe_2Mo_3O_{12}$, $Fe_2Mo_3O_8$, Fe_2MoO_4, and $FeMoO_4$ of the Fe－Mo－O Ternary System an μ Phase of the Fe－Mo Binary System by Electromotive Force Measurte Using a Y_2O_3-Stabilized ZrO_2 Solid Electrlyte K.Koyama, M.Morishita, T.Harada and N.Maekawa, Metallurgical and Materials Transactions, Vol. B34, (2003), pp. 653-659.	
$\Delta_f G^\circ_m$($FeMoO_4$(cr), 1112 - 1339 K) / kJ mol^{-1} = － 1053.5 + 0.2983 T ± 0.4	

Substance	$FeMoO_4$(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.	
$\Delta_{sln} G^\circ_m$($FeMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 46.73 ± 9.56	

Substance	$Fe_2Mo_3O_{12}$(cr)
Determination of Standard Gibbs Energies of Formation of $Fe_2Mo_3O_{12}$, $Fe_2Mo_3O_8$, Fe_2MoO_4, and $FeMoO_4$ of the Fe－Mo－O Ternary System an μ Phase of the Fe－Mo Binary System by Electromotive Force Measurte Using a Y_2O_3-Stabilized ZrO_2 Solid Electrlyte K.Koyama, M.Morishita, T.Harada and N.Maekawa, Metallurgical and Materials Transactions, Vol. B34, (2003), pp. 653-659.	
$\Delta_f G^\circ_m$($Fe_2Mo_3O_{12}$(cr), 1112 - 1339 K) / kJ mol^{-1} = － 2347 + 0.6814 T ± 1	

Substance	$Fe_2Mo_3O_8$(cr)
Determination of Standard Gibbs Energies of Formation of $Fe_2Mo_3O_{12}$, $Fe_2Mo_3O_8$, Fe_2MoO_4, and $FeMoO_4$ of the Fe－Mo－O Ternary System an μ Phase of the Fe－Mo Binary System by Electromotive Force Measurte Using a Y_2O_3-Stabilized ZrO_2 Solid Electrlyte K.Koyama, M.Morishita, T.Harada and N.Maekawa, Metallurgical and Materials Transactions, Vol.B34, (2003), pp.653-659.	
$\Delta_f G^\circ_m$($Fe_2Mo_3O_8$(cr), 1040 - 1145 K) / kJ mol^{-1} = － 2993 + 0.9105 T ± 2	

Substance	Fe_2MoO_4(cr)
Determination of Standard Gibbs Energies of Formation of $Fe_2Mo_3O_{12}$, $Fe_2Mo_3O_8$, Fe_2MoO_4, and $FeMoO_4$ of the Fe－Mo－O Ternary System an μ Phase of the Fe－Mo Binary System by Electromotive Force Measurte Using a Y_2O_3-Stabilized ZrO_2 Solid Electrlyte K.Koyama, M.Morishita, T.Harada and N.Maekawa, Metallurgical and Materials Transactions, B34, 653-659 (2003).	
$\Delta_f G^\circ_m$(Fe_2MoO_4(cr), 1243 - 1466 K) / kJ mol^{-1} = － 1174 + 0.342 T ± 1	

Substance $FeWO_4$(cr)
Determination of Standard Gibbs Energiers of Formation of $FeWO_4$ and Fe_2WO_8 by Electromotive Force Measurement, K.Koyama and T.Harada, J. Japan Soc. Powder Powder Metallurgy, Vo.40 (1993), pp.401-405.
$\Delta_f G^\circ_m$($FeWO_4$(cr), T K) / J mol^{-1} = − 118530 + 357.6 T ±200

Substance Fe_2WO_6(cr)
Determination of Standard Gibbs Energiers of Formation of $FeWO_4$ and Fe_2WO_8 by Electromotive Force Measurement, K.Koyama and T.Harada, J. Japan Soc. Powder Powder Metallurgy, Vo.40 (1993), pp.401-405.
$\Delta_f G^\circ_m$(Fe_2WO_6(cr), T K) / J mol^{-1} = − 1651900 + 516.1 T ±300

Substance $MgMoO_4$(cr)
Thermodynamic Properties for $MMoO_4$ (M = Mg, Sr and Ba) as the End-members of the Yellow Phases Formed in the Nuclear Fuel Waste Glasses, M. Morishita, Y. Kinoshita, A. Nozaki and H. Yamamoto Appl. Geo-chem., **98**, (2018),pp.310-320.
S°_m($MgMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 119.11 ± 1.19 $\Delta_f G^\circ_m$($MgMoO_4$(cr), 298.15 K) / kJ mol^{-1} = − 1295.73 ± 0.91 $\Delta_{sln} G^\circ_m$($MgMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 3.72 ± 1.88

Substance Mo(cr)
Thermodynamic Properties for $Sm_2(MoO_4)_3$: Standard Entropy; Neel Temperature; Solubility Product, M. Morishita, Y. Kinoshita, H. Tanaka, A. Nozaki and H. Yamamoto Monatshefte für Chemie - Chemical Monthly, **149**, (2018), pp.341-356.
$C^\circ_{p,m}$(Mo(cr), 298.15 K) / J K^{-1} mol^{-1} = 23.911 ± 0.057 S°_m(Mo(cr), 298.15 K) / J K^{-1} mol^{-1} = 28.573 ± 0.086

Substance MoO_3(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.
S°_m(MoO_3(cr), 298.15 K) / J K^{-1} mol^{-1} = 75.43 ± 0.75 $\Delta_f G^\circ_m$(MoO_3(cr), 298.15 K) / kJ mol^{-1} = −667.20 ± 0.63

Substance MoO_4^{2-}(aq)
Thermodynamic Properties of Molybdate Ion: Reaction Cycles and Experiments, H. Gamsjäger and M.Morishita, Pure and Applied Chemistry, **87** (2015) 461-476.
S°_m(MoO_4^{2-}(aq), 298.15 K) / J K^{-1} mol^{-1} = 32.03 ± 4.05 $\Delta_f H^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 996.807 ± 0.826 $\Delta_f G^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 836.542 ± 0.881

Substance MoO_4^{2-}(aq)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.
S°_m(MoO_4^{2-}(aq), 298.15 K) / J K^{-1} mol^{-1} = 32.25 ± 4.41 $\Delta_f H^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 996.81 ± 0.83 $\Delta_f G^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 836.61 ± 1.02 E°(MoO_4^{2-}(aq), 298.15 K) / V = 4.34 ± 0.01

Substance MoO_4^{2-}(aq)
Thermodynamic Properties for $MMoO_4$ (M = Mg, Sr and Ba) as the End-members of the Yellow Phases Formed in the Nuclear Fuel Waste Glasses, M. Morishita, Y. Kinoshita, A. Nozaki and H. Yamamoto Appl. Geo-chem., **98**, (2018),pp.310-320.
S°_m(MoO_4^{2-}(aq), 298.15 K) / J K^{-1} mol^{-1} = 32.30 ± 4.28 $\Delta_f H^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 996.81 ± 0.83 $\Delta_f G^\circ_m$(MoO_4^{2-}(aq), 298.15 K) / kJ mol^{-1} = − 836.63 ± 0.97 E°(MoO_4^{2-}(aq), 298.15 K) / V = 4.34 ± 0.01

Substance　　$Nd_2(MoO_4)_3$(cr)
Thermodynamic Properties for Neodymium Molybdate, Y. Kinoshita, M. Morishita, A. Nozaki and H.Yamamoto, J. Jpn. Inst. Met., **60**, (2019), pp.111-120.
S°_m($Nd_2(MoO_4)_3$(cr), 298.15 K) / J K^{-1} mol^{-1} = 409.38 ± 4.39 $\Delta_f G^\circ_m$($Nd_2(MoO_4)_3$(cr), 298.15 K) / kJ mol^{-1} = − 4072.1 ± 2.7 T_N ($Nd_2(MoO_4)_3$(cr)) / K = 1.43 ± 0.14 $\Delta_{sln} G^\circ_m$($Nd_2(MoO_4)_3$(cr), 298.15 K) / (kJ (mol of compd.)$^{-1}$) = 219.1 ± 13.7

Substance　　$Nd_2(MoO_4)_3$(cr)
Thermodynamic Properties for Neodymium Molybdate, Y. Kinoshita, M. Morishita, A. Nozaki and H.Yamamoto, Mater. Trans., **60**, (2019), pp.111-120.
S°_m($Nd_2(MoO_4)_3$(cr), 298.15 K) / J K^{-1} mol^{-1} = 439.29 ± 4.39 $\Delta_f G^\circ_m$($Nd_2(MoO_4)_3$(cr), 298.15 K) / kJ mol^{-1} = − 4072.13 ± 6.14 T_N ($Nd_2(MoO_4)_3$(cr)) / K = 1.55 ± 0.16 $\Delta_{sln} G^\circ_m$($Nd_2(MoO_4)_3$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 73.12 ± 2.33

Substance　　$NiMoO_4$(cr)
Calorimetric Study of Nickel Molybdate: Heat Capacity, Enthalpy and Gibbs Energy of Formation, M.Morishita and A.Navrotsky, Journal of the American Ceramic Society, Vol.86, (2003), Vol.1927-1932.
S°_m(α-$NiMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 118.0 ± 1.18 $\Delta_f H^\circ_m$(α-$NiMoO_4$(cr), 298.15 K) / kJ mol^{-1} = − 1026.0 ± 1.20 $\Delta_f G^\circ_m$(α-$NiMoO_4$(cr), 298.15 K) / kJ mol^{-1} = − 921.2 ± 1.25 $\Delta_f G^\circ_m$(α-$NiMoO_4$(cr), 800 - 1000 K) / kJ (mol of atoms)$^{-1}$ = − 1052.0 + 0.3712 T (± 7) $\Delta_f G^\circ_m$(β-$NiMoO_4$(cr), 1000 - 1380 K) / kJ (mol of atoms)$^{-1}$ = − 993.0 +0.3113 T (± 7)

Substance　　$NiMoO_4$(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.
$\Delta_{sln} G^\circ_m$($NiMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 38.82 ± 1.79

Substance $Sm_2(MoO_4)_3$(cr)
Thermodynamic Properties for $Sm_2(MoO_4)_3$: Standard Entropy; Néel Temperature; Solubility Product, M. Morishita, Y. Kinoshita, H. Tanaka, A. Nozaki and H. Yamamoto Monatshefte für Chemie - Chemical Monthly, **149**, (2018), pp.341-356.
S°_{m}($Sm_2(MoO_4)_3$(cr), 298.15 K) / J K^{-1} mol^{-1} = 400.14 ± 4.00 $\Delta_f G^{\circ}_{m}$($Sm_2(MoO_4)_3$(cr), 298.15 K) / kJ mol^{-1} = − 4048.71 ± 4.45 T_N ($Sm_2(MoO_4)_3$(cr)) / K = 1.30 ± 0.30 $\Delta_{sln} G^{\circ}_{m}$($Sm_2(MoO_4)_3$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 68.56 ± 1.88

Substance $SrMoO_4$(cr)
Thermodynamic Properties of Molybdate ion: Reaction Cycles and Experiments, H. Gamjäger and M.Morishita, Pure and Applied Chemistry, **87**, (2015), pp. 461-476.
$\Delta_{sln} G^{\circ}_{m}$($SrMoO_4$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 45.00 ± 1.40

Substance $SrMoO_4$(cr)
The Third Law Entropy of Strontium Molybdates, M. Morishita and H. Houshiyama, Materials Transactions, Vol.56 (2015), pp.545-549.
S°_{m}($SrMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 136.56 ± 1.37 Θ_D($SrMoO_4$(cr)) / K = 373 ± 6 $\Delta_f G^{\circ}_{m}$($SrMoO_4$(cr), 298.15 K) / kJ mol^{-1} = − 1437.05 ± 3.60

Substance $SrMoO_4$(cr)
Thermodynamic Properties for $MMoO_4$ (M = Mg, Sr and Ba) as the End-members of the Yellow Phases Formed in the Nuclear Fuel Waste Glasses, M. Morishita, Y. Kinoshita, A. Nozaki and H. Yamamoto Appl. Geo-chem., **98**, (2018),pp.310-320.
S°_{m}($SrMoO_4$(cr), 298.15 K) / J K^{-1} mol^{-1} = 136.44 ± 1.36 $\Delta_f G^{\circ}_{m}$($SrMoO_4$(cr), 298.15 K) / kJ mol^{-1} = − 1441.32 ± 1.36

Substance $ThMo_2O_8$(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.
$\Delta_{sln} G^{\circ}_{m}$($ThMo_2O_8$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 92.42 ± 21.24

Substance	$UMoO_6$(cr)
Thermodynamic Properties for Calcium Molybdate, Molybdenum Tri-Oxide and Aqueous Molybdate Ion, M. Morishita, Y. Kinoshita, H. Houshiyama, A. Nozaki and H. Yamamoto J. Chemical Thermodynamics, **114**, (2017), pp.30-43.	
$\Delta_{\mathrm{sln}}G^{\circ}_{\mathrm{m}}$($UMoO_6$(cr), 298.15 K) / (kJ (mol of MoO_4^{2-}(aq)$^{-1}$) = 68.33 ± 34.47	